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41651338.096774 93352749.058824 20841943.010204 29736666.510638 60550526.333333 18256482.864198 13440562.75 1526086112 39323816665 116518619880 94962035984 59956852836 95586027378 34916328106 76690730.090909 107064537888 13766280.705882 9286365.0571429 817201629.5 16727003.4 71673018417
2257071.5294118 28082708.361111 17178592.26087 53468765650 32927840308 11749316.703704 80649950290 49894301680

Relative Atomic Mass, Relative Molecular Mass and Relative Formula Mass

1. Calculate the relative formula mass for the following compounds

- a. NaCl
- b. H₂SO₄
- c. Al₂Cl₃
- d. Mg(OH)₂
- e. Sn(NO₃)₂
- f. Mg₃(PO₄)₂

2. Calculate the following relative molecular masses and then use the data to identify the molecule

- a. 35.5 x 2 =
- b. (1.0 x 2) + 16 =
- c. 12 + 16 =
- d. (1.0 x 2) + 32.1 + (16 x 4) =
- e. 14 + (1.0 x 3) =

3. Calculate the relative atomic mass of the following. Give your answers to 3 significant figures.

- a. Chlorine consists of 75.77% chlorine-35 and 24.23% chlorine-37.
- b. Europium consists of 47.8% Eu-151 and 52.2% Eu-153
- c. Antimony (Sb) is 57.2% Sb-121 and 42.8% Sb-123
- d. Copper consists of 69.1% Cu-63 and 30.9% Cu-65
- e. Thallium consists of 29.5% thallium-203 and 70.5% thallium-205
- f. Magnesium consists of 78.6% ²⁴Mg, 10.1% ²⁵Mg and 11.3% of ²⁶Mg
- g. Rubidium consists of the stable isotopes: ⁸⁵Rb (72.2%) and ⁸⁷Rb (27.8%).
- h. Analysis of a fifty pence piece showed up two isotopes of an unknown metal as outlined below

	Isotope 1	Isotope 2
Relative isotopic mass	63	65
% abundance	77.2	22.8

Calculate the Relative atomic mass of the metal and then suggest which metal this is.

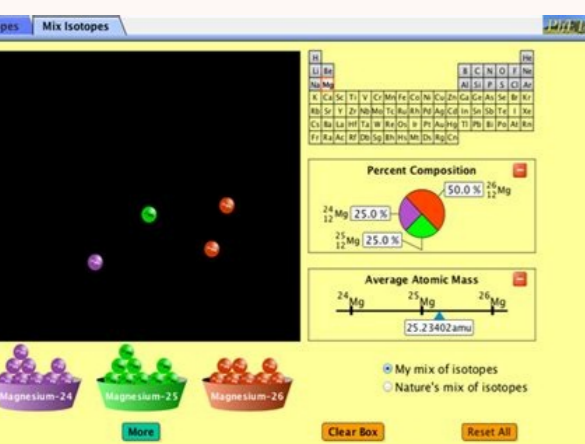
Worked Example 1 **Relative Atomic Mass**

Chlorine has two isotopes, ³⁵Cl and ³⁷Cl. The relative abundances of these isotopes are 75.77% and 24.23% respectively. Calculate the relative atomic mass of chlorine.

Isotope	Relative Abundance (%)	Relative Isotopic Mass
³⁵ Cl	75.77	35.46
³⁷ Cl	24.23	37.47

Worked Example 2 **Relative Atomic Mass**

Europium has two isotopes, ¹⁵¹Eu and ¹⁵³Eu. The relative abundances of these isotopes are 47.8% and 52.2% respectively. Calculate the relative atomic mass of europium.



Learning objectives define the atomic weight to calculate the atomic weight from the percentage of abundance manipulate the equation of the atomic weight to calculate variables variables Unknown distinguish between attempted weight, not up to the same as a mass of mass as mentioned in the previous section, which have the same attempt of prumens and nãutrons) are called ison (nucleus). There are ison and is the naturally occurring artificially produced. Of all the elements of the Perion Table, only 21 are pure elements. The pure or monotic elements are those elements with only one naturally occurs. The following lists the 21 pure elements: Table of Elements 2.3.1 Monotic isotopes of a given element not all exist in equal proportions. Mercyth, for example, has seven naturally: ¹⁹⁶Hg, ¹⁹⁸Hg, ¹⁹⁹Hg, ²⁰⁰Hg, ²⁰¹Hg, ²⁰²Hg, ²⁰⁴Hg. They are the percentage of natural abundance of 0.146%, 10.02%, 16.84%, 23.13%, 13.22%, 29.80%and 6.85%, respectively. It is clear that ²⁰²Hg occurs with greater abundance, and ²⁰⁰Hg is the most abundant near, but the other exists occur only in small locks. NOTE: The sum of the percentage of natural abundance of all ision of any element must total 100%. Some naturally and artificially produced radioactive artificial ision. All the heavier ãthomos However, there are many lighter nucienals than radioactive. For example, hydrogen has two isoable is the naturally occurring, ¹H and ²H (Deut), and a third ision Radioactive that occurs naturally, ³H (TREATY). It should not be surprising, but the ISOTETãPicas (% of each snow) poдем variar enmimboras. Agui estã; The Relatant Relatist of the country' This, ¶e how do we know what the percentage of abundance for each of the ishanakes of a particular element? The separated ision is through mass spectrometry; The BS tractions show the relative abundance of ision vs. No. Although we can not directly measure the mass of ãtomes, we can use the mass spectrum, an instrument that allows us to measure mass proposal to charge. In Figure 2.3.2, you can see the chlorine of the chlorine entering a mass spectrum. The chlorine has morals and is reached with a stream of ionizing trons that break the league of the cl2 and removes the chlorine trons that cause the formation of omes. These are accelerated in the cross -marary to come to a magnetic field that deviates the particles. The deflection of deflex depends on the mass of the partiam and the magnhanic field forcenial, with the clearest particles being diverted more (the lighter ³⁵Cl+ are diverted more than the heavier ³⁷Cl+). At the end of the mara is an output orifier with a detector, and as the intensity of the magnetic field increases, the deflex shift is changed, which separates the particles. Note that the mass spectrum in Figure 2.3.2 (B) provides the relative abundance of each ision, with the peak normalized for the ision with the largest abundance. Therefore, if this proposal was 3: 1, it means that there are 3 ³⁵Cl parts for each ³⁷Cl particle, and the percentage abundance would be 75% ³⁵Cl and 25% ³⁷Cl. Figure 2.3.2 Determination of relative to -the -mic masses using a dough -meter below is a youtube video that describes the mass spectrum here is a bars showing abundan Relative 4 ision of estrons Example: The strict mass spectrum has four different peaks, varying in intensity. The four peaks indicate that there are four estrons of strict. The four of strontium have isotopic mass numbers of 84, 86, 87 and 88, and relative abundance of 0.56%, 0.56%, 7.00% and 82.58% respectively, the intensity of the peak corresponds to abundance. ⁸⁴Sr has the lowest peak, which corresponds to its relative abundance of 0.56% to what ⁸⁸Sr has the highest peak, which corresponds to its relative abundance of 8.58% that indicates that ⁸⁸Sr is the isotope that occurs in higher quantities. Since we collect the relative masses of each isotope of mass spectrometry data, we can pray this information to calculate the average atomic mass (weight) of all atoms of an element taking into account the mass of each present isotope and the abundance percent for each isotope. this can be done through the following formula: average atomic mass = (mass of isotope 1 x fractional abundance of isotope 1) + (mass of isotope 2 x fractional abundance of isotope 2) + . the average atomic mass was calculated in this way and can be found under each symbol in the periodic table. Let us see an example of how we can calculate this information. average calculation of atomic mass problem 1 average of atomic mass: what is the average atomic mass of neon, since it has 3 isotopes with the abundances of follow-up percentage; ²⁰Ne = 19.992 amu (90.51),% ²¹Ne = 20.993 amu (0.27),% ²²Ne = 21.991 amu. what we know: once you know what the element is, you can solve it without doing math by using the periodic table, but you need to be able to do mathematics because it can be an unknown, and this is the only way you can discover the correct meaningful figures. since Ne-20 has the highest abundance percent, it must have the greatest impact on its average. Therefore, we hope that the average atomic mass is closer to the mass of Ne-20 (about 19.992 amu), click the following video tutor to see if we estimate correctly, video tutor: answer: according to the correct number of significant figures, we arrived with 20.18 amu as the average atomic weight until thought that the oxiba adartnocne e orole ed somot;ã ed sopot'Asi so sodot ed assam ed aid@ãm a ©ã ocim' ãta osep O jc atsopseR 53; 7254.53 jd 71; 7254.53 jc 53; 71 jb 71; 53 ja. ©ã 53-orole ed ocim' ãta orem'ãn o e ©ã aniroc ad ocim' ãta osep O ã) 2{ xednlepa ((oicãcrexE. aicneãdnuba ed roiam megatnecrop amu met e 1 opot'Asi o euq siam otium aid©ãm a uotcapmi 2 opot'Asi o euq acidni ossl.)UMA 885.921(1 opot'Asi od assam a euq aid@ãm ad amix'ãrp siam jãtse juma219.131(2 opot'Asi od assam a euq ecepaP. UMA 442.131 ©ã aid@ãm A. ossid s;ãrt rop oipãcnrp o ecehnoc 'ãcov es amelborp essen opmet otium ratsag ed eadidseeen jãh oeãn. 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General Quimmic; Principles and Modern Application. Ninth ed. New Jersey: Pearson Prentice Hall, 2007. HOUSECRAFT, Catherine E. and Alan G. Sharpe. Inorgain quamic. Third ed. England: Pearson Prentice Hall, 2008. HOEFS, Jochen. Esion Geoquamic Stop. Friday ed. Germany: Springer, 2009 Anonymous modified by Joshua Halpern, Scott Sinex and Scott Johnson Bob Belford and November Palmer Ronia Kattoum (Uair) In this spreadsheet, we will practice calculating the percentage of isotorical abundance of relative mass and the isotic masses. Q1: Chlorine has two stable ision, ³⁵Cl and ³⁷Cl, with athanic pasta 34.9689 U and 36.9659 U, respectively. The relative abundance of ³⁷Cl in a sample of chlorine is 3735cl = 0.3196. Calculate the absolute abundance of ³⁷Cl day. A46.97% B19.02% C51.56% D24.22% E40.49% Calculate the mass of chlorine. A35.61 U B36.48 U C36.33 U D35.97 U E35.45 U Q2: GUIZ HAS two stable ision, ⁶⁹Ga and ⁷¹Ga, with athanic masses 68.9256 U and 70.9247 U, respectively. Calculate, for 3 significant numbers, the 71ga abundance for a giving sample with a mass at the time 69,723 u. Calculate, for 3 significant numbers, the abundance of 69ga for a sample of guallo with a mass of 69,620 u. Q3: The magnã ©ã sio has TRONS ISMAIBE, 24mg, 25mg and 26mg, with athanic masses 23.9850 U, 24.9858 U and 25.9826 U, respectively. At gm42 gm42 ed siarutan 25Mg are 78.99% and 10.00% respectively. Calculate the average atomic mass of magnesium. A24.31 u B24.98 u C24.20 u D24.10 u E24.88 u The relative abundances of magnesium isotopes in a sample are 2524MgMg=0.11815 and 2624MgMg=0.14687. Calculate, to 4 significant figures, the absolute abundance of ²⁴Mg in the sample. A97.21% B58.10% C79.05% D77.98% E73.50% Q4: Antimony has two stable isotopes, ¹²¹Sb and ¹²³Sb, with atomic masses 120.9038 u and 122.9042 u respectively. The relative abundance of ¹²³Sb in an average sample of antimony is 123121SbSb=0.7479. Calculate the average absolute abundance of ¹²³Sb. A42.79%ããã B14.42%ããã C33.71%ããã D20.13%ããã E25.21%ããã Calculate the average atomic mass of antimony. A121.86 u B122.05 u C121.90 u D121.41 u E121.76 u Q5: Silicon has three stable isotopes, ²⁸Si, ²⁹Si, and ³⁰Si, with atomic masses 27.9769 u, 28.9765 u, and 29.9738 u respectively. The natural abundances of ²⁸Si and ²⁹Si are 92.23% and 4.68%

respectively. Calculate, to 4 significant figures, the average atomic mass of silicon. A28.05 u B28.03 u C28.09 u D28.11 u E28.02 u The relative abundances of silicon isotopes in a sample are 2928SiSi=0.08069 and 3029SiSi=0.49583. Calculate, to 4 significant figures, the absolute abundance of 28Si in the sample. A63.43% B63.10% C88.97% D89.23% E87.93% E87.93%

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